

**Directions:** Choose the letter of the correct answer.

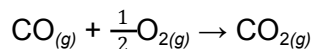
Consider the reaction shown below.



MM: CO = 28 g/mol; O<sub>2</sub> = 32 g/mol; CO<sub>2</sub> = 44 g/mol

- Which of the following statements is TRUE?
  - The reaction is heat-releasing.
  - The reaction absorbs heat from the surroundings.
  - The reverse process is exothermic.
  - None of the above

- What is the corresponding  $\Delta H$  for the reaction below?



- 283 kJ/mol
  - 566 kJ/mol
  - +566 kJ/mol
  - 1132 kJ/mol
- One of the factors below is incorrect. Which is it?
    - $\frac{-566 \text{ kJ}}{2 \text{ mol CO}_{(g)}}$
    - $\frac{-566 \text{ kJ}}{56 \text{ g CO}_{(g)}}$
    - $\frac{-566 \text{ kJ}}{2 \text{ mol O}_{2(g)}}$



## Thermochemistry

### Practice Questions

- d.  $\frac{-566 \text{ kJ}}{32 \text{ g } O_{2(g)}}$
4. In the presence of excess  $O_2$ , how much heat will be liberated if 98 g of CO will undergo the reaction?
- 990.5 kJ
  - 630.2 kJ
  - 1981.0 kJ
  - 1734.3 kJ
5. If the reaction produces 1.5 moles of  $CO_2$  before depletion of the limiting reactant, how much heat was liberated?
- 212.3 kJ
  - 424.5 kJ
  - 849.0 kJ
  - 1698.0 kJ



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*To God be the glory!*